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The empirical formula for a compound is C_2H_5 and its relative formula mass is 58.0. Deduce its molecular formula.
Deduce its molecular formula. (Relative atomic masses: $C = 12.0$, $H = 1.0$)

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g/mole. b. Empirical formula is CH₂Cl₂,
mass is 99.0 g/mole. c. Empirical formula
is C₃H₄ mass is 121 g/ mole. stop i HSPI -
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Empirical Formula Pogil Science Division Bellevue empirical formula. 2. Divide the molecular mass by the empirical mass. 3.

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Multiply all of the subscripts in the empirical formula by the resulting factor.
Example Problem: What is the molecular formula for a material that is 92.2% C and 7.80% H by mass? The material has a molar mass of 52.0 g/mol. 1. Molecular mass /

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